

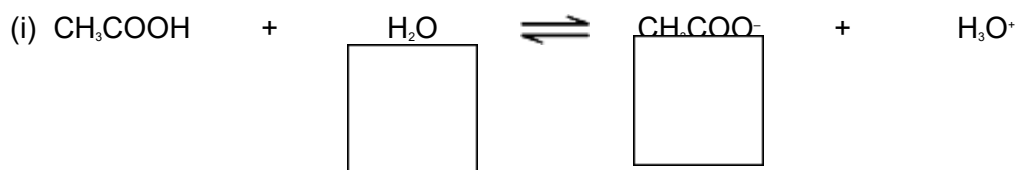
Q1. This question is about several Brønsted–Lowry acids and bases.

(a) Define the term *Brønsted–Lowry acid*.

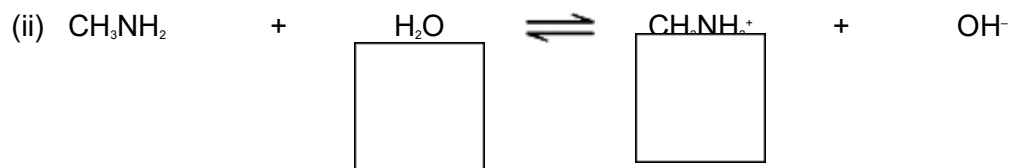
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(1)

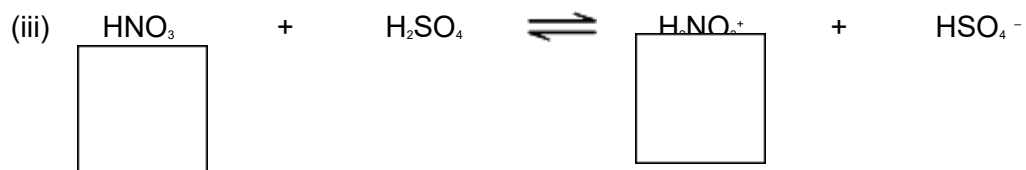
(b) Three equilibria are shown below. For each reaction, indicate whether the substance immediately **above** the box is acting as a Brønsted–Lowry acid (**A**) or a Brønsted–Lowry base (**B**) by writing **A** or **B** in each of the six boxes.



(1)



(1)



(1)

(c) A 25.0 cm³ sample of 0.0850 mol dm⁻³ hydrochloric acid was placed in a beaker. Distilled water was added until the pH of the solution was 1.25.

Calculate the total volume of the solution formed. State the units.

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(Extra space)
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(3)

(d) At 298 K, the value of the acid dissociation constant (K_a) for the weak acid HX in aqueous solution is $3.01 \times 10^{-5} \text{ mol dm}^{-3}$.

(i) Calculate the value of $\text{p}K_a$ for HX at this temperature.
Give your answer to 2 decimal places.

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(1)

(ii) Write an expression for the acid dissociation constant (K_a) for the weak acid HX.

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(1)

(iii) Calculate the pH of a $0.174 \text{ mol dm}^{-3}$ solution of HX at this temperature.
Give your answer to 2 decimal places.

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(6)

(Total 18 marks)

Q2. Ammonia and ethylamine are examples of weak Brønsted–Lowry bases.

(a) State the meaning of the term *Brønsted–Lowry base*.

.....
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(1)

(b) (i) Write an equation for the reaction of ethylamine ($\text{CH}_3\text{CH}_2\text{NH}_2$) with water to form a weakly alkaline solution.

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(1)

(ii) In terms of this reaction, state why the solution formed is **weakly** alkaline.

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.....

(1)

(c) State which is the stronger base, ammonia or ethylamine. Explain your answer.

Stronger base

Explanation

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.....
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(Extra space)

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(3)

(d) Give the formula of an organic compound that forms an alkaline buffer solution when added to a solution of ethylamine.

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(1)

(e) Explain qualitatively how the buffer solution in part (d) maintains an almost constant pH when a small amount of hydrochloric acid is added to it.

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(Extra space)

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(2)

(Total 9 marks)

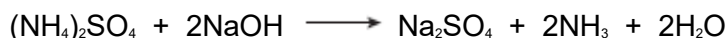
Q3.When iron(II) sulfate is used for killing weeds in lawns, it is often mixed with the fertiliser ammonium sulfate. Ammonium sulfate also makes the soil acidic.

(a) Write an equation to show how the ammonium ion behaves as a Brønsted–Lowry acid in water.

.....

(1)

- (b) Compounds such as ammonium sulfate react on warming with sodium hydroxide solution as shown in the equation below.



Use this information to describe a simple test, other than smell, to show that ammonia is evolved. State what you would observe.

Test

.....

Observation

(2)
(Total 3 marks)

Q4. In which one of the following reactions is the role of the reagent stated correctly?

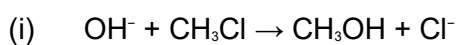
	Reaction	Role of reagent
A	$\text{TiO}_2 + 2\text{C} + 2\text{Cl}_2 \rightarrow \text{TiCl}_4 + 2\text{CO}$	TiO_2 is an oxidising agent
B	$\text{HNO}_3 + \text{H}_2\text{SO}_4 \rightarrow \text{H}_2\text{NO}_3^+ + \text{HSO}_4^-$	HNO_3 is a Brønsted-Lowry acid
C	$\text{CH}_3\text{COCl} + \text{AlCl}_3 \rightarrow \text{CH}_3\text{CO}^+ + \text{AlCl}_4^-$	AlCl_3 is a Lewis base
D	$2\text{CO} + 2\text{NO} \rightarrow 2\text{CO}_2 + \text{N}_2$	CO is a reducing agent

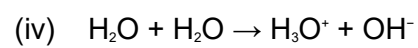
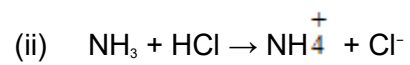
(Total 1 mark)

Q5.

Summarised directions for recording responses to multiple completion questions			
A (i), (ii) and (iii) correct only	B (i) and (iii) correct only	C (ii) and (iv) correct only	D (iv) alone correct

Brønsted-Lowry acid-base reactions include





(Total 1 mark)