

Q1. In this question, give all values of pH to 2 decimal places.

(a) The ionic product of water has the symbol K_w

(i) Write an expression for the ionic product of water.

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(1)

(ii) At 42°C, the value of K_w is $3.46 \times 10^{-14} \text{ mol}^2 \text{ dm}^{-6}$.

Calculate the pH of pure water at this temperature.

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(iii) At 75 °C, a $0.0470 \text{ mol dm}^{-3}$ solution of sodium hydroxide has a pH of 11.36. Calculate a value for K_w at this temperature.

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(b) Methanoic acid (HCOOH) dissociates slightly in aqueous solution.

(i) Write an equation for this dissociation.

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(ii) Write an expression for the acid dissociation constant K_a for methanoic acid.

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(1)

- (iii) The value of K_a for methanoic acid is $1.78 \times 10^{-4} \text{ mol dm}^{-3}$ at 25°C . Calculate the pH of a $0.0560 \text{ mol dm}^{-3}$ solution of methanoic acid.

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- (iv) The dissociation of methanoic acid in aqueous solution is endothermic.

Deduce whether the pH of a solution of methanoic acid will increase, decrease or stay the same if the solution is heated. Explain your answer.

Effect on pH

Explanation

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(3)

- (c) The value of K_a for methanoic acid is $1.78 \times 10^{-4} \text{ mol dm}^{-3}$ at 25°C . A buffer solution is prepared containing $2.35 \times 10^{-2} \text{ mol}$ of methanoic acid and $1.84 \times 10^{-2} \text{ mol}$ of sodium methanoate in 1.00 dm^3 of solution.

- (i) Calculate the pH of this buffer solution at 25°C .

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- (ii) A 5.00 cm³ sample of 0.100 mol dm⁻³ hydrochloric acid is added to the buffer solution in part (c)(i).

Calculate the pH of the buffer solution after this addition.

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(4)
(Total 20 marks)

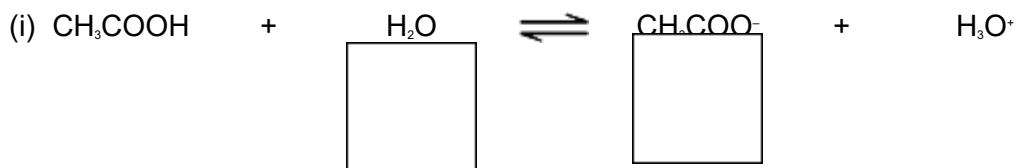
Q2. This question is about several Brønsted–Lowry acids and bases.

- (a) Define the term *Brønsted–Lowry acid*.

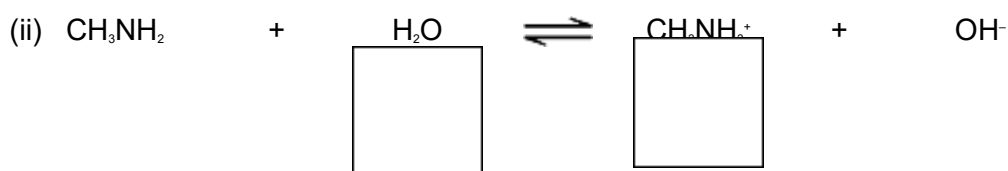
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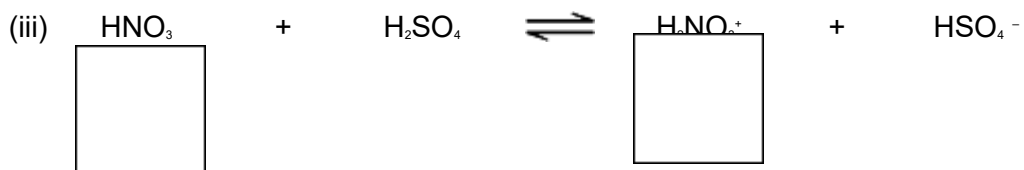
(b) Three equilibria are shown below. For each reaction, indicate whether the substance immediately **above** the box is acting as a Brønsted–Lowry acid (**A**) or a Brønsted–Lowry base (**B**) by writing **A** or **B** in each of the six boxes.



(1)



(1)



(1)

(c) A 25.0 cm³ sample of 0.0850 mol dm⁻³ hydrochloric acid was placed in a beaker. Distilled water was added until the pH of the solution was 1.25.

Calculate the total volume of the solution formed. State the units.

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(d) At 298 K, the value of the acid dissociation constant (K_a) for the weak acid HX in aqueous solution is $3.01 \times 10^{-5} \text{ mol dm}^{-3}$.

(i) Calculate the value of $\text{p}K_a$ for HX at this temperature.
Give your answer to 2 decimal places.

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(1)

(ii) Write an expression for the acid dissociation constant (K_a) for the weak acid HX.

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(1)

(iii) Calculate the pH of a $0.174 \text{ mol dm}^{-3}$ solution of HX at this temperature.
Give your answer to 2 decimal places.

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(3)

- (e) An acidic buffer solution is formed when 10.0 cm^3 of $0.125 \text{ mol dm}^{-3}$ aqueous sodium hydroxide are added to 15.0 cm^3 of $0.174 \text{ mol dm}^{-3}$ aqueous HX. The value of K_a for the weak acid HX is $3.01 \times 10^{-5} \text{ mol dm}^{-3}$.

Calculate the pH of this buffer solution at 298 K.
Give your answer to 2 decimal places.

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(Total 18 marks)

Q3.A student was given a task to determine the percentage purity of a sample of salicylic acid. The method used by the student to prepare a solution of salicylic acid is described below.

- 0.500 g of an impure sample of salicylic acid was placed in a weighing bottle.
- The contents were tipped into a beaker and 100 cm³ of distilled water were added.
- Salicylic acid does not dissolve well in cold water so the beaker and its contents were heated gently until all the solid had dissolved.
- The solution was poured into a 250 cm³ graduated flask and made up to the mark with distilled water.

(a) Give **two** additional instructions that would improve this method for making up the salicylic acid solution.

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(2)

(b) The pH of this solution was measured and a value of 2.50 was obtained. Calculate the concentration of salicylic acid in this solution. Assume that salicylic acid is the only acid in this solution. The K_a for salicylic acid is $1.07 \times 10^{-3} \text{ mol dm}^{-3}$. You may represent salicylic acid as HA. Show your working.

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(3)

(c) Use your answer to part (b) to calculate the mass of salicylic acid ($M_r = 138.0$) present in the original sample. (If you were unable to complete the calculation in part (b), assume that the

concentration of salicylic acid is $8.50 \times 10^{-3} \text{ mol dm}^{-3}$. This is **not** the correct answer.)

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(2)

- (d) Use your answer to part (c) to calculate the percentage purity of the salicylic acid used to make the solution.
(If you were unable to complete the calculation in part (c), assume that the mass of salicylic acid is 0.347 g. This is **not** the correct answer.)

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(1)

(Total 8 marks)