

Please write clearly in block capitals.

Centre number

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Candidate number

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Surname

Forename(s)

Candidate signature

I declare this is my own work.

GCSE COMBINED SCIENCE: TRILOGY

F

Foundation Tier
Chemistry Paper 2F

Time allowed: 1 hour 15 minutes

Materials

For this paper you must have:

- a ruler
- a scientific calculator
- the periodic table (enclosed).

Instructions

- Use **black ink** or **black ball-point pen**.
- Pencil should only be used for drawing.
- **Fill** in the boxes at the top of this page.
- Answer **all** questions in the spaces provided.
- If you need extra space for your answer(s), use the lined pages at the end of this book. Write the question number against your answer(s).
- Do **all** rough work in this book. Cross through any work you do not want to be marked.
- In **all** calculations, show clearly how you work out your answer.

Information

- The maximum mark for this paper is 70.
- The marks for questions are shown in brackets.
- You are expected to use a calculator where appropriate.
- You are reminded of the need for good English and clear presentation in your answers.

For Examiner's Use	
Question	Mark
1	
2	
3	
4	
5	
6	
7	
TOTAL	

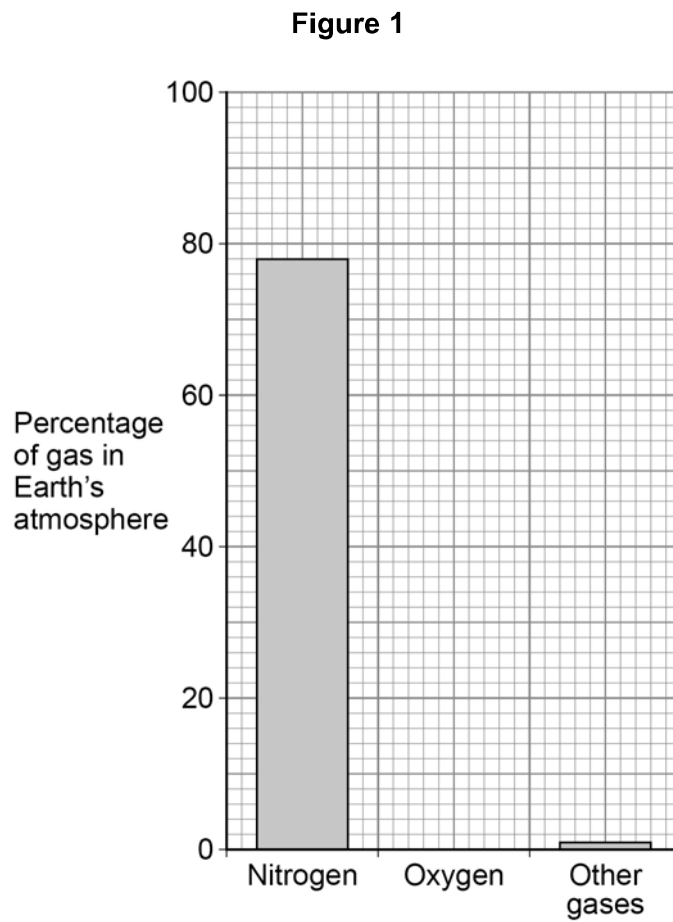


0 1

This question is about the Earth's atmosphere.

0 1 . 1

The Earth's atmosphere contains 21% oxygen.

Draw the bar for oxygen on **Figure 1**.**[1 mark]**

0 1 . 2

What is used to test for oxygen gas?

[1 mark]Tick (✓) **one** box.

A burning splint

A glowing splint

Damp litmus paper

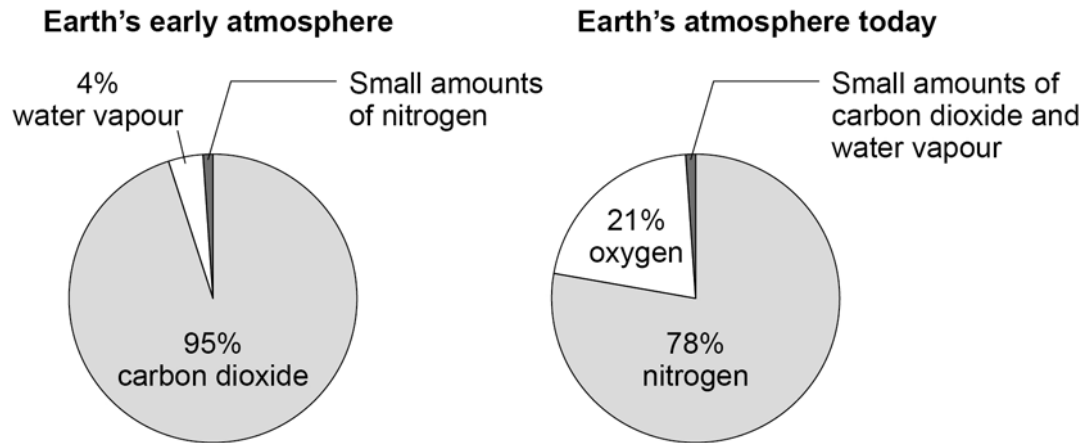
Limewater



The Earth's early atmosphere was very different from the Earth's atmosphere today.

Figure 2 shows the composition of the Earth's early atmosphere and of the Earth's atmosphere today.

Figure 2



0 1 3

The percentages of nitrogen and oxygen in the Earth's atmosphere today are different from the Earth's early atmosphere.

Complete the sentences.

Choose answers from the box.

Use **Figure 2**.

Each answer can be used once, more than once or not at all.

[2 marks]

decreased increased stayed the same

Since the Earth's early atmosphere, the percentage of nitrogen in the Earth's atmosphere has _____.

Since the Earth's early atmosphere, the percentage of oxygen in the Earth's atmosphere has _____.

Turn over ►



0 1 . 4

The Earth's atmosphere today contains a small amount of carbon dioxide.

Why has the percentage of carbon dioxide decreased since the Earth's early atmosphere?

[2 marks]

Tick (✓) **two** boxes.

Dissolved in oceans

Formation of sedimentary rocks

Industrialisation

Respiration

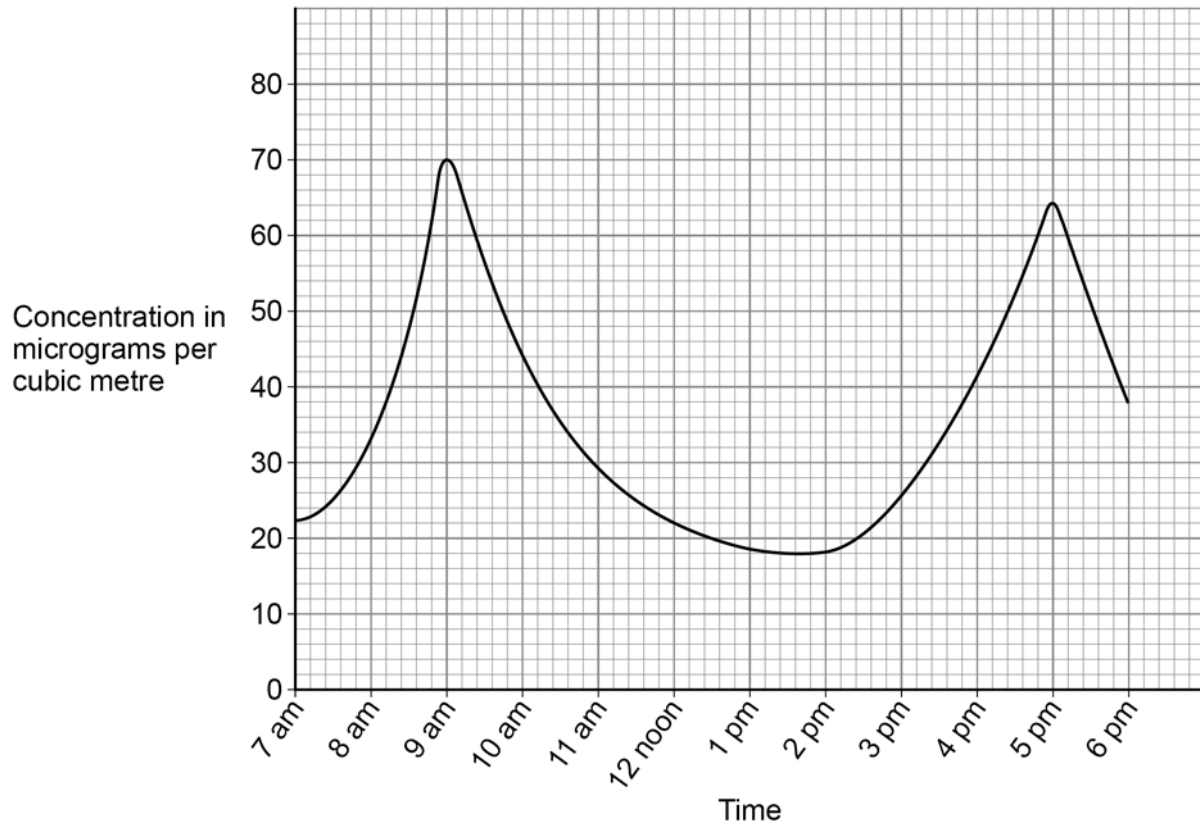
Volcanic activity



Oxides of nitrogen are produced when nitrogen reacts with oxygen in car engines.

Figure 3 shows the concentration of oxides of nitrogen in the atmosphere during one day in a city.

Figure 3



0 1 . 5

Which **two times** have the highest concentrations of oxides of nitrogen in the atmosphere?

[2 marks]

1 _____

2 _____

0 1 . 6

Suggest why there are the highest concentrations of oxides of nitrogen at these times.

[1 mark]



0 2

This question is about fuels.

Coal deposits were formed from the remains of trees.

0 2 . 1

Name the process in the leaves of trees that uses carbon dioxide.

[1 mark]

0 2 . 2

How is coal formed after trees die?

[1 mark]

Tick (✓) **one** box.

The trees are burned.

The trees are compressed.

The trees are melted.

Coal contains small amounts of sulfur.

0 2 . 3

Name the gas produced when sulfur burns in oxygen.

[1 mark]

0 2 . 4

Give **two** problems caused by the gas produced when sulfur burns in oxygen.

[2 marks]

1 _____

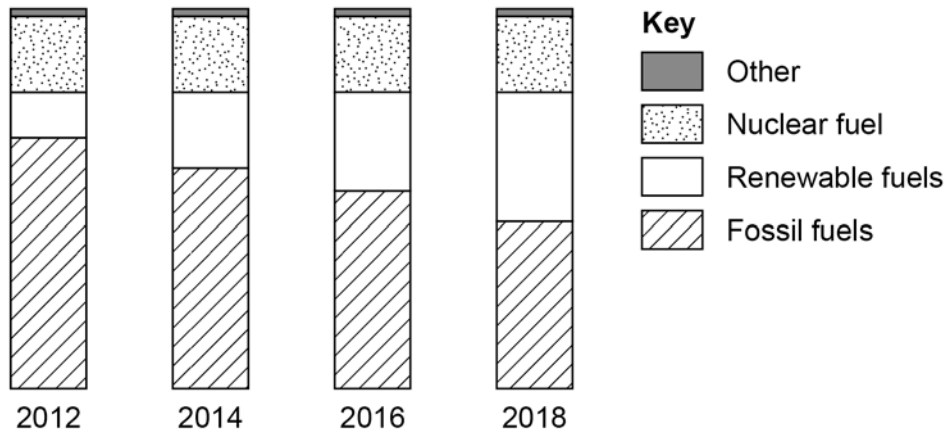
2 _____



0 2 . 5

Figure 4 shows the relative amount of electricity generated from different fuel sources in the UK from 2012 to 2018.

Figure 4



Describe what happens to the amounts of fuels used to generate electricity in the UK from 2012 to 2018.

[3 marks]

8

Turn over for the next question

Turn over ►

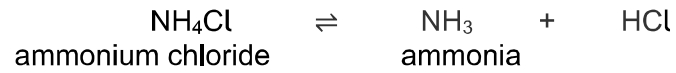


0 3

This question is about ammonia and its compounds.

A student heated a sample of ammonium chloride.

The equation for the reaction is:



0 3 . 1

One product is ammonia.

What is the name of the product with the formula HCl?

[1 mark]

0 3 . 2

Ammonia is a gas.

What is the state symbol for ammonia?

[1 mark]

Tick (✓) **one** box.

(aq) (g) (l) (s)

0 3 . 3

How does the equation show that the reaction is reversible?

[1 mark]

0 3 . 4

Complete the sentence.

[1 mark]

The forward reaction is endothermic,

so the reverse reaction is _____.



0 3 . 5 Complete the sentence.

Choose the answer from the box.

[1 mark]

concentration

rate

temperature

Equilibrium is reached when the forward and reverse reactions happen at exactly the same _____ .

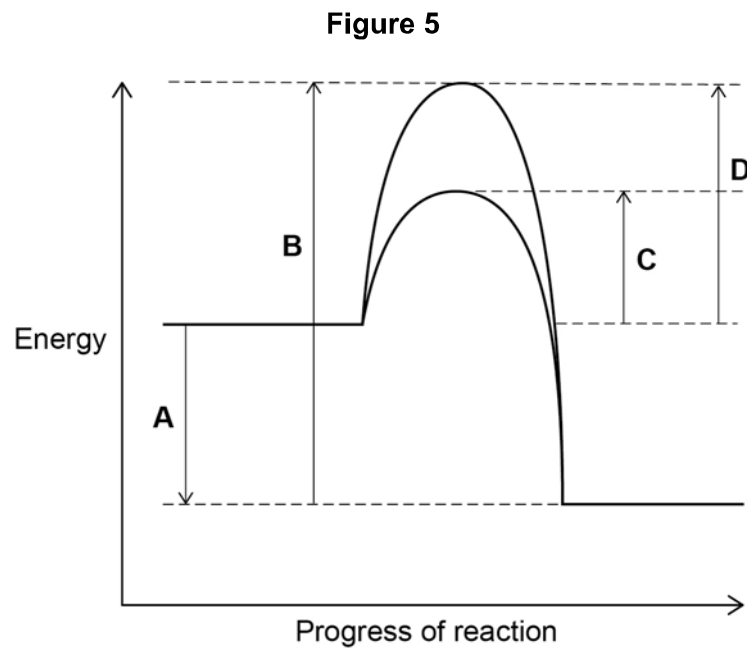
Question 3 continues on the next page

Turn over ►



The industrial process to produce ammonia uses a catalyst.

0 3 . 6 **Figure 5** shows the reaction profile for the reaction with and without a catalyst.



Which letter represents the activation energy for the reaction with a catalyst?

[1 mark]

Tick (✓) **one** box.

A B C D



0 3 . 7 Give **one** reason why using a catalyst reduces costs.

Do **not** answer in terms of activation energy.

[1 mark]

0 3 . 8 Ammonia is in a mixture that is used as a household cleaner.

What is a mixture that has been designed as a useful product called?

[1 mark]

8

Turn over for the next question

Turn over ►



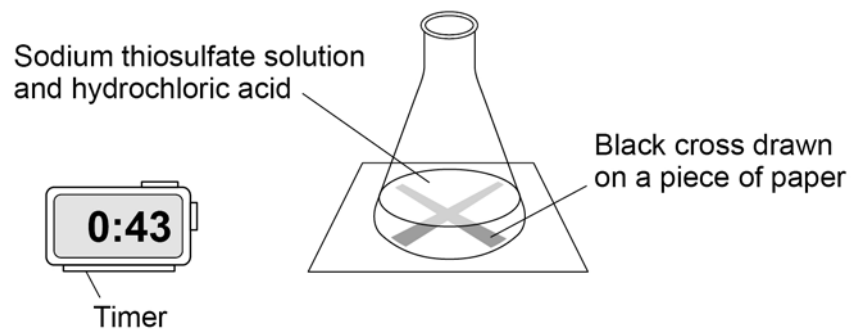
0 4

A student investigates the effect of concentration on the rate of the reaction between sodium thiosulfate solution and hydrochloric acid.

Figure 6 shows the experiment.

The experiment was done in a fume cupboard.

Figure 6



This is the method used.

1. Pour 50 cm³ of sodium thiosulfate solution into a conical flask.
2. Put the conical flask on a black cross drawn on a piece of paper.
3. Pour 10 cm³ of hydrochloric acid into the conical flask and start a timer.
4. Stop the timer when the cross can no longer be seen.
5. Repeat the experiment with different concentrations of sodium thiosulfate solution.



0 4 . 1

Draw **one** line from each type of variable to the correct example of the variable in this investigation.

[2 marks]**Type of variable****Example of variable**

Dependent

Concentration of sodium
thiosulfate solutionTemperature of reaction
mixtureTime taken for the cross
to no longer be seen

Independent

Volume of acid

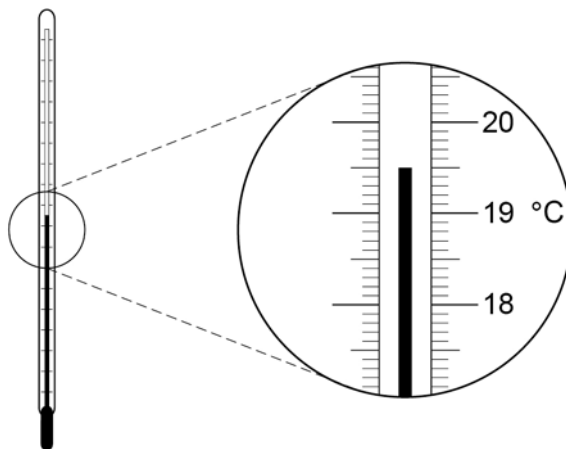
Volume of the flask

Question 4 continues on the next page**Turn over ►**

0 4 . 2

The experiment is done at room temperature.

Figure 7

What is the temperature shown on the thermometer in **Figure 7**?**[1 mark]**

Temperature = _____ °C

Table 1 shows the student's results.**Table 1**

Concentration of sodium thiosulfate solution in mol/dm ³	Time in seconds
0.1	82
0.2	40
0.3	20
0.4	13
0.5	10
0.6	8

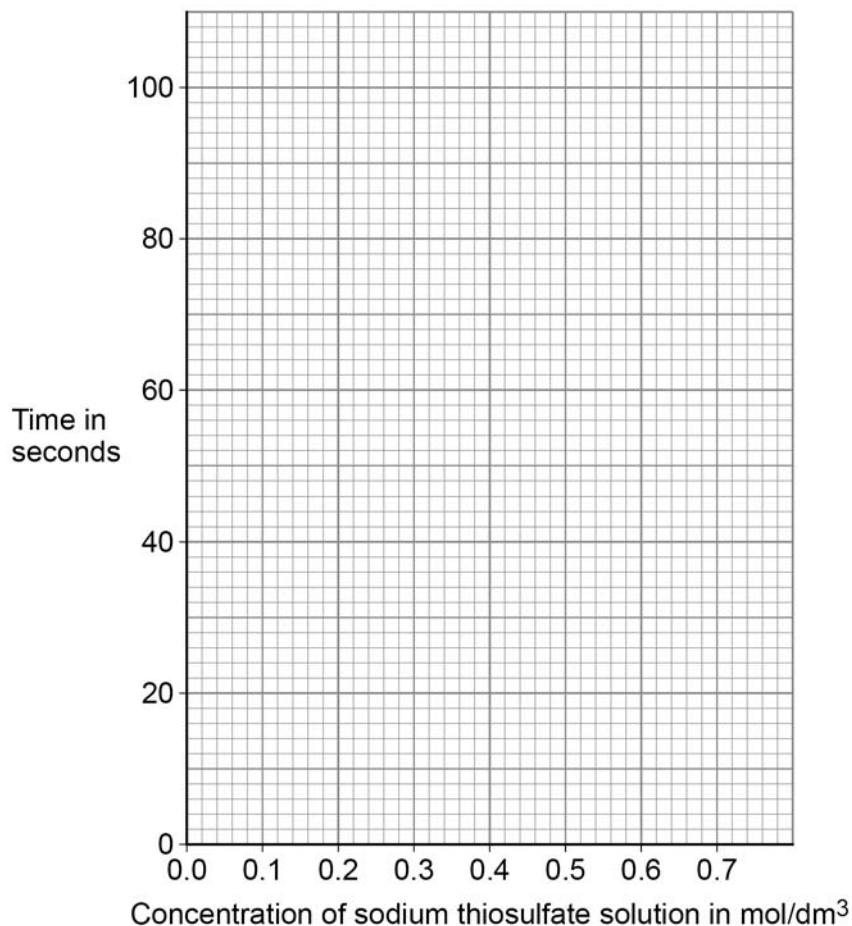


0 4 . 3 Plot the data from **Table 1** on **Figure 8**.

Draw a line of best fit.

[3 marks]

Figure 8



0 4 . 4 Predict the time taken for the cross to no longer be seen at a concentration of 0.7 mol/dm^3

Use your graph in **Figure 8**.

[1 mark]

Time = _____ s

0 4 . 5 Complete the sentence.

[1 mark]

As the concentration of sodium thiosulfate solution increases, the time taken for the cross to no longer be seen _____ .

Turn over ►



0 4 . 6

In one experiment 0.725 g of sulfur is produced in 20 seconds.

Calculate the mean rate of the reaction from 0 to 20 seconds.

Use the equation:

$$\text{mean rate of reaction} = \frac{\text{mass of sulfur produced in grams}}{\text{time in seconds}}$$

[2 marks]

Mean rate of reaction = _____

0 4 . 7What is the unit for the mean rate of reaction calculated in Question **04.6**?**[1 mark]**Tick (✓) **one** box.g g/s s s/g 

0 4 . 8

The student did the experiment with 0.15 mol/dm^3 sodium thiosulfate solution and repeated the experiment three more times.

Table 2 shows the results.

Table 2

	Test 1	Test 2	Test 3	Test 4
Time in seconds for the cross to no longer be seen	60.5	63.2	82.3	65.7

Calculate the mean time for this reaction.

Do **not** include the anomalous result in your calculation.

Give your answer to 3 significant figures.

[3 marks]

Mean time for the reaction (3 significant figures) = _____ s

14

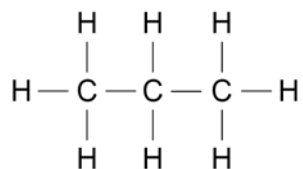
Turn over for the next question

Turn over ►



0 5

This question is about hydrocarbons.

Figure 9 shows a hydrocarbon.**Figure 9**

0 5 . 1

Complete the formula for the hydrocarbon shown in **Figure 9**.

[1 mark]

C _____ H _____

0 5 . 2

What is the name of the hydrocarbon in **Figure 9**?

[1 mark]

0 5 . 3

Which homologous series does the hydrocarbon in **Figure 9** belong to?

[1 mark]



0 5 . 4 30 g of another hydrocarbon contains 24 g of carbon.

Which calculation gives the percentage of carbon in the hydrocarbon?

[1 mark]

Tick (✓) **one** box.

$$\frac{24 \times 30}{100} \quad \square$$

$$\frac{100 \times 30}{24} \quad \square$$

$$\frac{24 \times 100}{30} \quad \square$$

$$\frac{24}{30 \times 100} \quad \square$$

0 5 . 5 **Table 3** shows boiling points of some hydrocarbons.

Table 3

Formula of hydrocarbon	Boiling point in °C
C ₂ H ₆	-89
C ₄ H ₁₀	0
C ₆ H ₁₄	69
C ₈ H ₁₈	125
C ₁₀ H ₂₂	174

Describe how the boiling points change as the number of carbon atoms in the hydrocarbon increases.

[1 mark]

Turn over ►



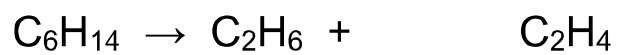
Hydrocarbons can be cracked.

0 5 . 6 Give **one** condition used to crack hydrocarbons.

[1 mark]

0 5 . 7 Balance the equation for the cracking of C_6H_{14}

[1 mark]



0 5 . 8 Give **one** reason why hydrocarbons are cracked.

[1 mark]



0 5 9

Window frames can be manufactured from wood or plastic.

Table 4 shows the results of a life cycle assessment (LCA) for making one wooden and one plastic window frame.

Both window frames are the same size.

Table 4

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Give **three** advantages of using wood instead of plastic in the manufacture of window frames.

[3 marks]

Advantage of wood 1 _____

Advantage of wood 2 _____

Advantage of wood 3 _____

11

Turn over for the next question

Turn over ►



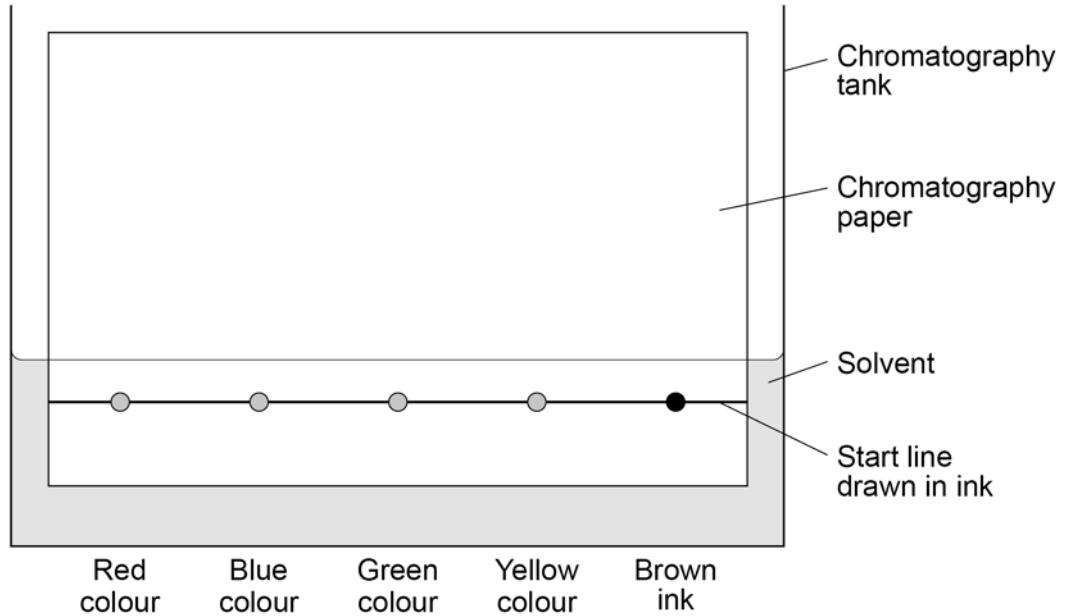
0 6

A student investigated the colours in a brown ink using chromatography.

0 6 . 1

Figure 10 shows the apparatus used.

Figure 10



Give **two** errors made by the student.

Describe the problem each error would cause.

[4 marks]

Error 1 _____

Problem 1 _____

Error 2 _____

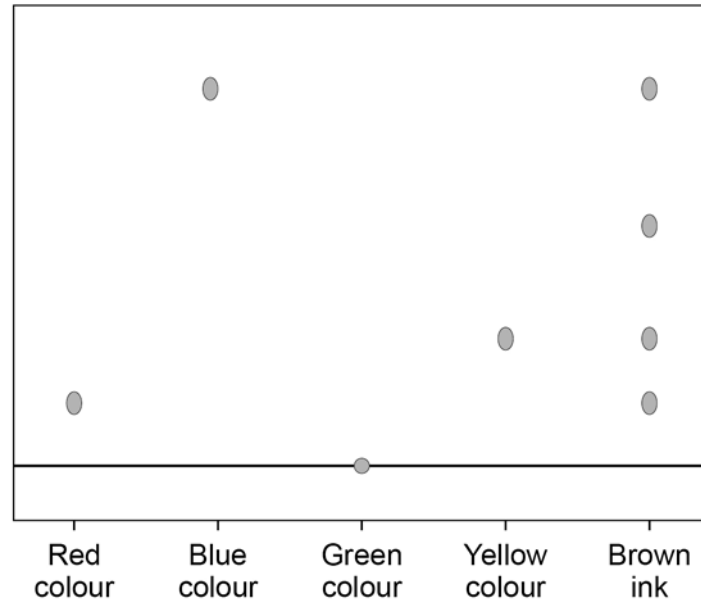
Problem 2 _____



A different student set up the apparatus correctly.

Figure 11 shows the results.

Figure 11



0 6 . 2

Give **two** conclusions the student can make from **Figure 11** about the four colours in the brown ink.

[2 marks]

1 _____

2 _____

Question 6 continues on the next page

Turn over ►



0 6 . 3 Why was the green colour still on the start line at the end of the experiment?

[1 mark]

Tick (✓) **one** box.

The experiment was left for too long.

The green colour was insoluble in the solvent.

The green spot contained too many colours.

The green spot was too small.

0 6 . 4 A student calculated the R_f value of a colour to be 0.24

The colour moved 1.8 cm from the start line.

Calculate the distance the solvent moved.

Use the equation:

$$R_f = \frac{\text{distance moved by colour}}{\text{distance moved by solvent}}$$

[3 marks]

Distance moved by solvent = _____ cm

10



0 7 . 1

Water that is safe to drink is called potable water.

Compare how easily potable water can be obtained from:

- waste water (sewage)
- ground water (fresh water).

[6 marks]

Question 7 continues on the next page

Turn over ►



A scientist produced potable water from 150 cm^3 of salty water.

0 7 . 2 Which process can be used to produce potable water from salty water?

[1 mark]

Tick (✓) **one** box.

Distillation

Electrolysis

Filtration

Sterilisation

0 7 . 3 The salty water contains sodium chloride.

The scientist collected 2.40 g of sodium chloride from 150 cm^3 of salty water.

Calculate the concentration of sodium chloride in grams per dm^3

[3 marks]

Concentration of sodium chloride = _____ g/dm^3

10

END OF QUESTIONS



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3 2



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