

Mark Scheme (Results)

Summer 2013

International GCSE  
Chemistry (4CH0) Paper 2CR

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Publications Code UG035550

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Question number	Answer	Notes	Marks
1 (a)	gallium / Ga		1
(b)	sodium / magnesium / aluminium / Na / Mg / Al		1
(c)	fluorine / F / F <sub>2</sub>		1
(d)	nitrogen / N / N <sub>2</sub>		1
(e)	neon / argon / krypton / xenon / radon / Ne / Ar / Kr / Xe / Rn		1
		<b>Total</b>	<b>5</b>

Question number	Answer	Notes	Marks
2 (a)	B A D C		1 1 1 1
(b)	Mixture Compound Mixture		1 1 1
		<b>Total</b>	<b>7</b>

Question number	Answer	Notes	Marks
3 (a)	hydrogen / H <sub>2</sub> burns with a pop/squeak OR use burning/lit splint/flame to see if pop/squeak	Ignore H Must be reference to test and result Reference to splint/match with no indication of flame is not enough Reject reference to glowing splint Ignore flame extinguished 'Squeaky pop test' on its own is not sufficient	1 1
(b) i	AgCl  (dilute) nitric acid / HNO <sub>3</sub>	Ignore names even if wrong Accept sulfuric acid / H <sub>2</sub> SO <sub>4</sub> Reject hydrochloric acid / HCl Ignore concentrated acid Ignore acid(ified) without a named acid Reject other named acids	1 1
ii	iron nitrate	Accept ferrous nitrate and ferric nitrate Ignore oxidation states (II) and (III) Reject other oxidation states	1

Question number	Answer	Notes	Marks
3 (c)	(add) sodium hydroxide (solution) / NaOH  green precipitate  brown precipitate	<p>Any group I hydroxide / ammonium hydroxide / barium or calcium hydroxide / ammonia solution (names or formulae)            If reagent incorrect, then 0/3            If reagent missing, then M2 and M3 can be awarded            If near miss (eg ammonia hydroxide) then M2 and M3 can be awarded</p> <p>Ignore qualifiers such as light / pale / dark            Accept solid / suspension / ppt(e) in place of precipitate            Reject all other colours            Ignore names and formulae even if incorrect</p> <p>Ignore qualifiers such as light / pale / dark / rusty / foxy / orange            Accept red-brown            Accept solid / suspension / ppt(e) in place of precipitate            Reject all other colours            Ignore names and formulae even if incorrect</p> <p>If both colours correct, penalise missing precipitate once only            Do not award M2 or M3 for two correct observations in the wrong order            Ignore references to bubbles etc</p>	1  1  1
		<b>Total</b>	<b>8</b>

Question number	Answer	Notes	Marks
4 (a)	bubbles / fizzing / effervescence  sodium moves / darts / floats sodium gets smaller / disappears sodium melts / forms ball white trail	Accept gas given off/evolved/formed/produced Accept hydrogen gas Ignore identity of gas  Accept equivalents such as shoots/skims Accept dissolves  Do not apply list principle Assume that it = sodium Ignore flames / sparks Any two for 1 each	2
(b)	Do not apply list principle	Assume that it = sodium	1
(c) i	hydrogen / H <sub>2</sub>	Ignore H	1
ii	K <sup>+</sup>		1

Question number	Answer	Notes	Marks
4 (d)	<p>Na is 2.8.1 K is 2.8.8.1</p> <p>outer/valence electron / outer shell / electron lost in K further from nucleus/protons</p> <p>less attracted by nucleus</p>	<p>Accept other punctuation and no punctuation and diagrams in place of full stops If neither of M1 and M2 scored, allow potassium has more (electron) shells (or numbers of shells stated)/energy levels for 1 mark?</p> <p>Ignore potassium further from nucleus</p> <p>Accept (electron) more easily removed/lost /less energy needed to remove (electron) Accept potassium more willing to lose electron If no reference to nucleus or protons, then neither M3 nor M4 can be awarded A correct reference to nucleus/protons is needed before M3 and M4 can be awarded Ignore references to shielding Accept reverse arguments for sodium in M3 and M4</p>	<p>1 1</p> <p>1</p> <p>1</p>
		<b>Total</b>	<b>9</b>



Question number	Answer			Notes	Marks
5 (a)	<b>Statement</b>	<b>Fractional distillation</b>	<b>Cracking</b>	1 mark for each line correct	5
	Crude oil is heated	(✓)			
	A catalyst may be used		✓		
	Alkenes are formed		✓		
	Decomposition reactions occur		✓		
	Fuels are obtained	✓	✓		
	Separation is the main purpose	✓			
(b) i	$C_5H_{12}$			Accept $H_{12}C_5$	1
ii	<pre>       H H H H H                 H—C—C—C—C—C—H                       H H H H H </pre>				1
iii	$C_5H_{12}$				1
iv	pentane				1
v	$C_nH_{2n+2}$				Accept x and other letters in place of n Accept answers like $C_nH_{2n+2}$ Ignore $2(n+1)$

Question number	Answer	Notes	Marks
5 (c) i	(products) 2 2 (oxygen) 3	M1 and M2 independent	1 1
ii	4 electrons shared between 2 (carbon) atoms 4 electron pairs between 2C and 4H atoms	Ignore inner electrons even if wrong Ignore number of hydrogen atoms  Accept all permutations of dots and crosses Ignore intersecting circles Accept H atoms at all angles At least one C or one H atom must be labelled – max 1 if not Max 1 if more than 2 C atoms Max 1 if wrong number of electrons in outer shell of any atom	1 1
(d) i	phosphoric acid / H <sub>3</sub> PO <sub>4</sub> any value in range 250 – 350 °C	Ignore concentrated / dilute Accept value without unit Accept 523 – 623 <u>K</u> Marks independent	1 1
ii	20 (mol) M1 × 24 480 (dm <sup>3</sup> )	Accept 480 000 <u>cm</u> <sup>3</sup> If M1 incorrect but 480 is final answer, then only M3 can be awarded If no answer to amount of ethene, then 20 x 24 = 480 scores M2 and M3	1 1 1
		<b>Total</b>	<b>19</b>

Question number	Answer	Notes	Marks
6 (a)	ethanol/it is more volatile/evaporates more quickly/more easily/evaporates in a shorter time	Accept has a lower boiling point (than water) Ignore reference to melting point(s) Accept reverse arguments for water	1
(b) i	0.3(0) (g)	Accept some copper dropped off  Any two for 1 each	1
ii	some copper did not stick to (negative) electrode/cathode some copper removed during washing/drying <u>positive</u> electrode/anode impure OR formed (anode) sludge		2

Question number	Answer	Notes	Marks
6 (c) i	<p>all 9 points plotted correctly to nearest gridline</p> <p>straight line of best fit</p>	<p>Deduct 1 mark for each error Award these marks if points too faint to be seen under correct line Ignore point at 0.55</p> <p>Must be drawn with a ruler Must go through origin Ignore extrapolation beyond (16,0.5)</p>	
ii	point at (7.40, 0.20) circled		
iii	<p>no charge/current/electricity passed AND no copper deposited/no change in mass/no electrolysis</p>	<p>OWTTE, eg charge = 0, so mass (increase) = 0 Ignore references to direct proportion</p>	
iv	<p>line is straight / fixed gradient AND goes through origin</p>	<p>Ignore re-statements of the information given in the question, eg the greater the charge, the greater the mass (increase)</p>	
v	graph line extrapolated to (at least) 0.55 correct value from candidate graph	<p>Probably 17.4 - 17.8 M2 not dependent on extrapolation</p>	
		<b>Total</b>	<b>12</b>
		<b>Total for paper</b>	<b>60</b>



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Order Code UG035550 Summer 2013

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