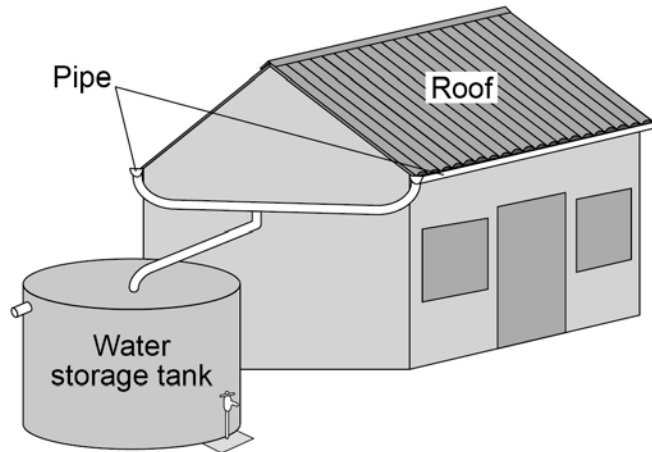


0 1

Rainwater is collected from the roofs of houses as shown in **Figure 1**.

Figure 1



0 1 . 1 The water in the storage tank is **not** potable.

What does potable mean?

Tick **one** box.

[1 mark]

- | | |
|-------------------------------|--------------------------|
| Contains dissolved substances | <input type="checkbox"/> |
| Pure | <input type="checkbox"/> |
| Safe to drink | <input type="checkbox"/> |
| Tastes nice | <input type="checkbox"/> |

0 1 . 2 Why should the water in the tank be filtered to make it potable?

[1 mark]

Tick **one** box.

To kill microbes

To remove dissolved gases

To remove dissolved solids

To remove undissolved solids

0 1 . 3 A gas which bleaches litmus paper can be added to the water to make it potable.

Name this gas and explain why it is added.

[2 marks]

Question 1 continues on the next page

The storage tank is made from concrete reinforced with steel wire, as shown in **Figure 2**.

Figure 2

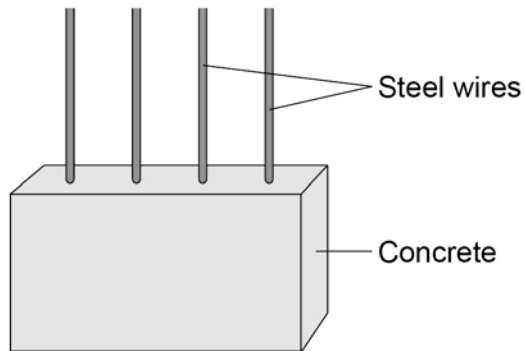
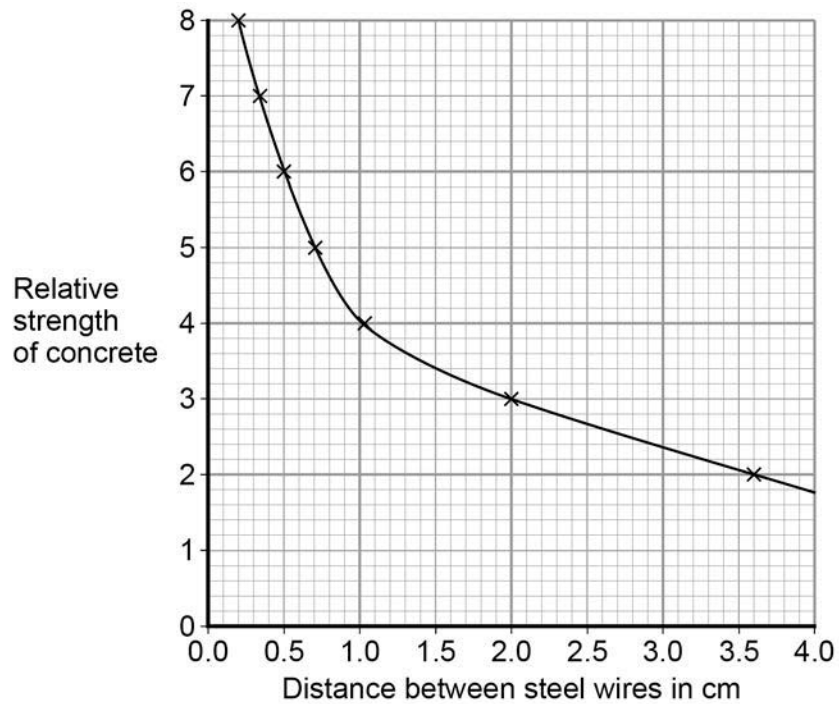


Figure 3 shows how the distance between the steel wires affects the relative strength of the concrete.

Figure 3



0	1	.	4
---	---	---	---

Use values from **Figure 3** to describe the relationship shown by the graph.

[2 marks]

Turn over for the next question

0 2 . 1 The hydrocarbon $C_{16}H_{34}$ can be cracked.

Balance the equation for cracking $C_{16}H_{34}$

[1 mark]



0 2 . 2 Describe the differences between cracking and distillation.

[2 marks]

0 2 . 3 What type of reaction is cracking?

[1 mark]

Tick **one** box.

Combustion

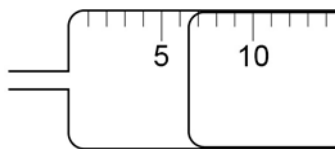
Decomposition

Neutralisation

Precipitation

0 3 . **2** **Figure 5** shows the gas syringe during one of the experiments.

Figure 5



What is the volume of gas collected?

[1 mark]

Tick **one** box.

5.3 cm³

6.0 cm³

6.5 cm³

7.0 cm³

Question 3 continues on the next page

0 3 . 3 Figure 6 shows the student's results for one concentration of hydrochloric acid.

Figure 6

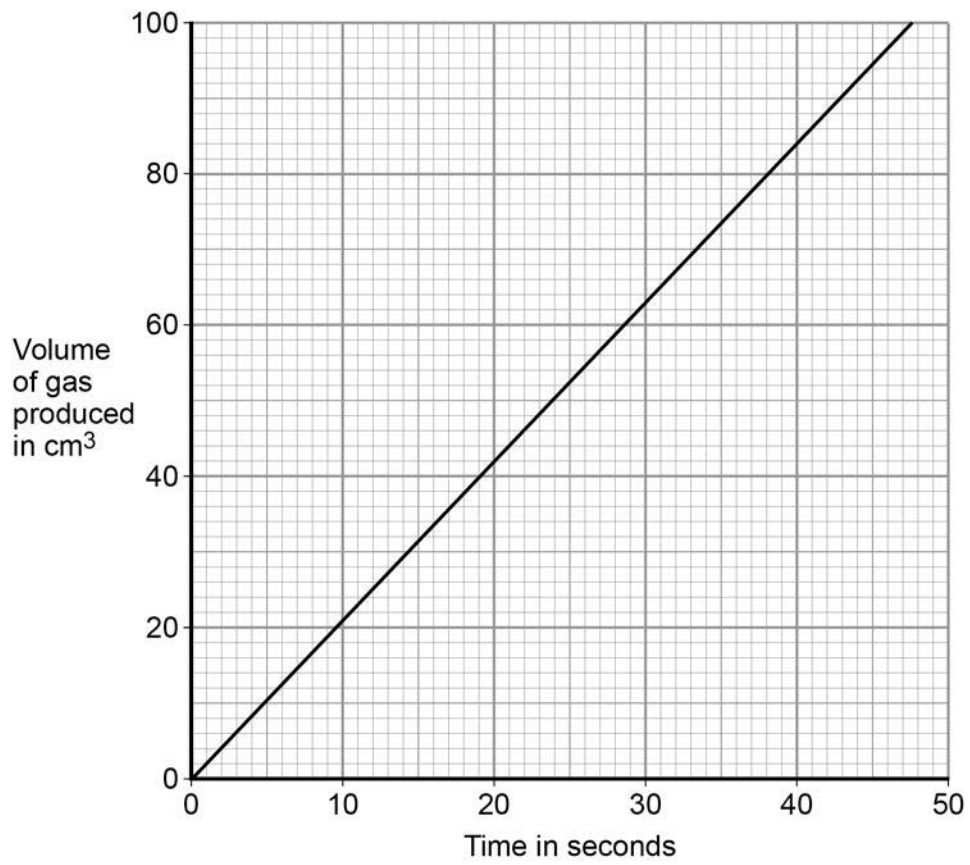


Table 2 shows the student's results when the concentration was two times greater than the results on **Figure 6**.

Table 2

Time in seconds	Volume of gas produced in cm ³
0	0
10	35
15	52
20	80
30	87

Plot the results in **Table 2** on the grid in **Figure 6**.
Draw a line of best fit.

[3 marks]

0 3 . 4

Give **one** conclusion about how the rate of reaction changed when the concentration of hydrochloric acid was changed.

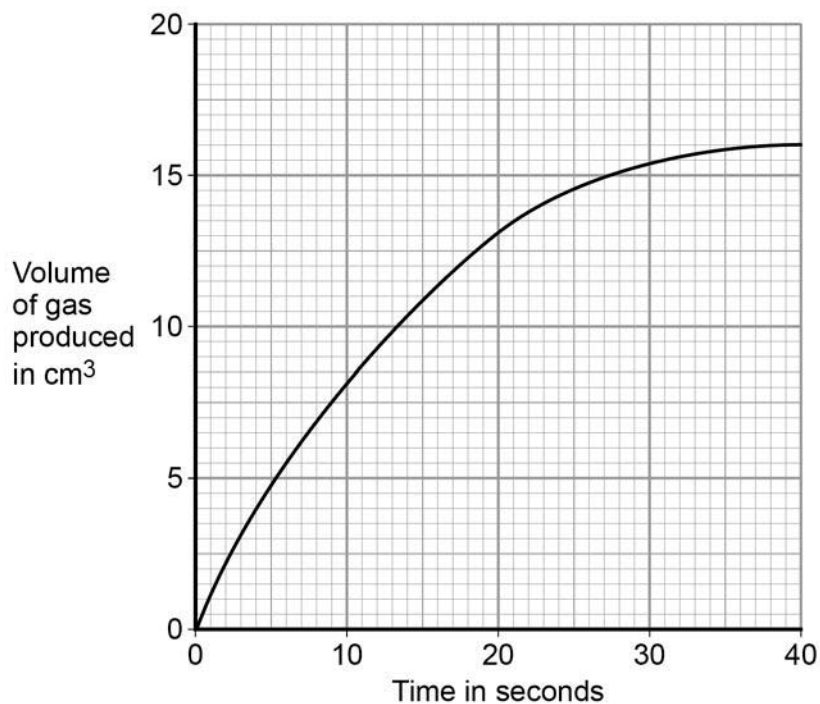
[1 mark]

Question 3 continues on the next page

0 3 . 5

Figure 7 shows volume of gas produced against time for the reaction between magnesium and ethanoic acid.

Figure 7



Draw a tangent to the curve at 20 seconds.

Determine the rate of the reaction at 20 seconds by calculating the gradient of the tangent.

Give the unit.

[4 marks]

Rate of reaction = _____

Unit = _____

0 3 . **6** Explain, in terms of particles, why the rate decreases during the reaction between magnesium and ethanoic acid.

[2 marks]

Turn over for the next question

0 4 Aqamed is a medicine for children.

0 4 . **1** The medicine is a formulation.

What is meant by a formulation?

[1 mark]

0 4 . **2** Children often do not like taking medicine.

Suggest a substance that could be added to Aqamed to increase the desire for children to take it.

Give a reason for your suggestion.

[2 marks]

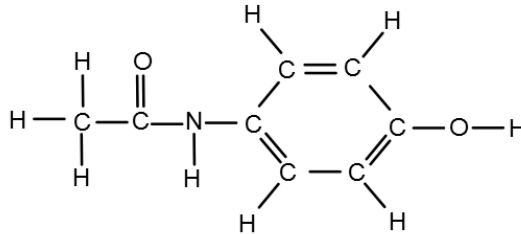
Substance _____

Reason _____

0 4 . 3 The main ingredient in Aqamed is a painkiller called paracetamol.

Figure 8 represents a molecule of paracetamol.

Figure 8



Give the molecular formula of paracetamol.

Calculate its relative formula mass (M_r).

Relative atomic masses (A_r): H = 1; C = 12; N = 14; O = 16

[2 marks]

Molecular formula _____

Relative formula mass _____

$M_r =$ _____

Question 4 continues on the next page

0 4 . 4 Aspirin is a medicine for use by adults.

An aspirin tablet contains 300 mg of acetylsalicylic acid.

Calculate the number of moles of acetylsalicylic acid in one aspirin tablet.

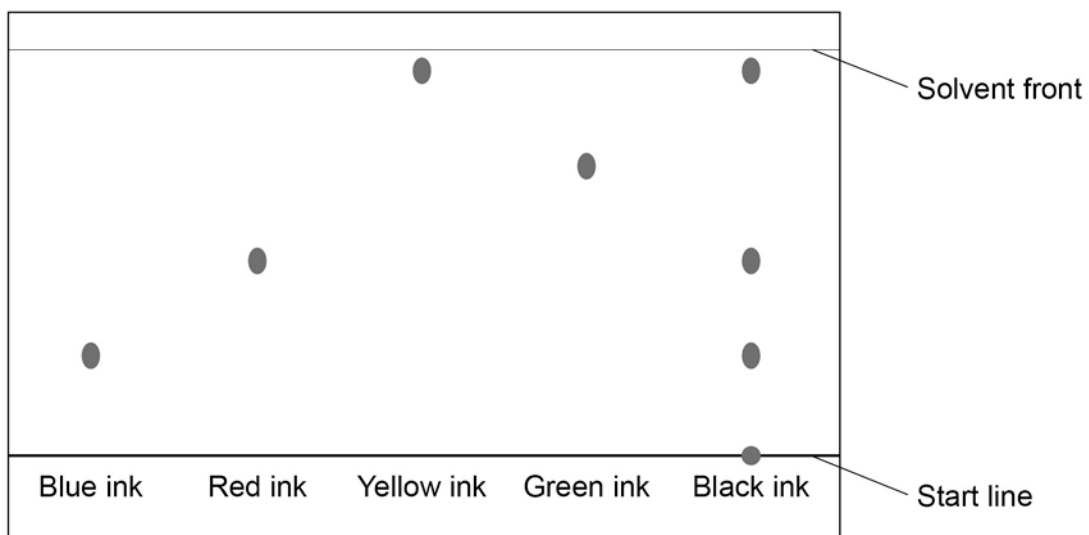
Give your answer in standard form to three significant figures.

Relative formula mass (M_r) of aspirin = 180

[4 marks]

Number of moles = _____

Turn over for the next question

0 5**Figure 9** shows a paper chromatogram of five different inks.**Figure 9****0 5****1** Explain how paper chromatography separates substances.**[3 marks]**

0 5 . 2

Analyse the chromatogram. Describe and explain the result for black ink.

[4 marks]

0 5 . 3

Use **Figure 9** to calculate the R_f value of the blue ink.

[3 marks]

R_f value = _____

Turn over for the next question

There are no questions printed on this page

0 6

There is less carbon dioxide in the Earth's atmosphere now than there was in the Earth's early atmosphere.

0 6 . **1**

The amount of carbon dioxide in the Earth's early atmosphere decreased because it was used by plants and algae for photosynthesis, dissolved in the oceans and formed fossil fuels.

Give **one** other way that the amount of carbon dioxide in the Earth's early atmosphere decreased.

[1 mark]

0 6 . **2**

Carbon dioxide is a greenhouse gas.

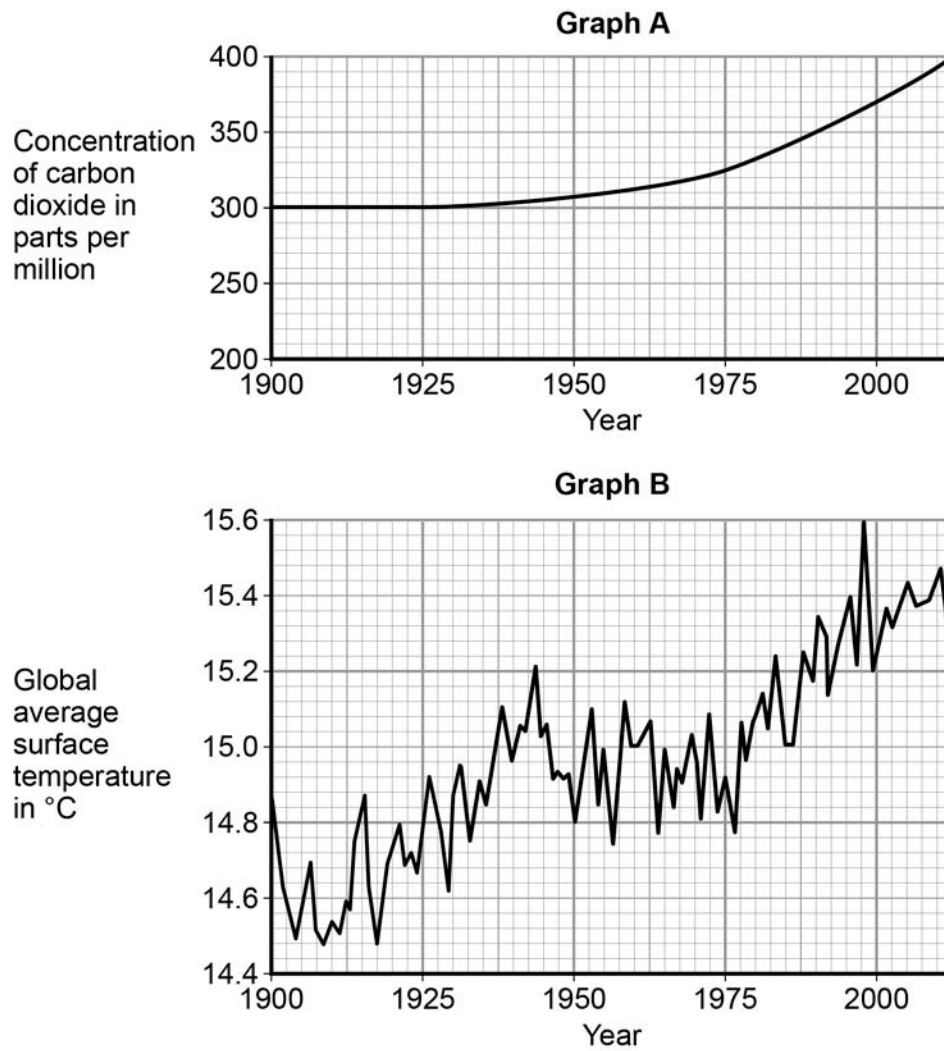
Describe the greenhouse effect.

[4 marks]

Question 6 continues on the next page

The graphs in **Figure 10** show the concentration of carbon dioxide in the atmosphere and global average surface temperature since 1900.

Figure 10



0 6 . **3** Calculate the percentage increase in the concentration of carbon dioxide from 1975 to 2000.

[1 mark]

_____ %

0 6 . **4** What was the global average surface temperature in 1980?

[1 mark]

Global average surface temperature = _____ °C

0 6 . **5** A student stated: 'The graphs show that increasing the concentration of carbon dioxide in the atmosphere causes global temperature increases.'

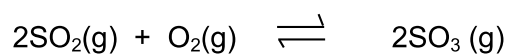
Discuss why this statement is only partially true.

[4 marks]

Turn over for the next question

0 7Sulfur dioxide (SO₂) is used to manufacture sulfuric acid.**0 7**. **1** Explain why sulfur dioxide has a low boiling point.**[3 marks]**

The equation shows one stage in the manufacture of sulfuric acid from sulfur dioxide.



The reaction is exothermic in the forward direction.

0 7. **2** Use Le Chatelier's Principle to predict the effect of increasing the temperature on the amount of sulfur trioxide (SO₃) produced at equilibrium.

Give a reason for your answer.

[2 marks]

0	7	.	3
---	---	---	---

Use Le Chatelier's Principle to predict the effect of increasing the pressure on the amount of sulfur trioxide (SO_3) produced at equilibrium.

Give a reason for your answer.

[2 marks]

END OF QUESTIONS

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